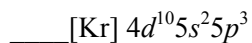
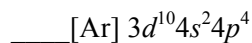
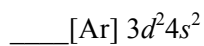
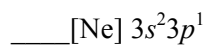


Chapter 5-Periodic Properties Worksheet

Give the symbol for the elements with the following configurations



Use the elements listed below to answer the following questions

K Ca Na Li Ar F Cl B C P Ti

____ Has the lowest ionization energy

____ Has the smallest atomic radius

____ Would take the largest amount of energy to remove an electron

____ Would most strongly attract electrons in a compound

____ Would need to absorb the most amount of energy to take on an electron

____ Is a noble gas

____ Is a transition metal

____ Would release the most energy when taking on an electron

____ Would need the least amount of energy to remove an electron

____ Is the largest atom

Of the pairs below pick which larger

____ Na or Na^+

____ Li or F

____ F or F^-

____ H or H^-

____ Ar or Kr

____ Ti or Hf

____ Li or Fr

____ Ca or Ca^{+2}

____ Si or Pb

____ He or H

How does nuclear charge effect electronegativity, atomic radius, and ionization energy?